

Periodic Table of Elements

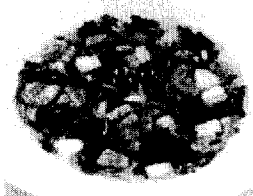
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37	38	39	40								48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90
91	92																			101	102	103	104	105	106	107	108	109																									

It All Matters Review

1. What is an element? Find an element in the room that can be seen and identified. Check the periodic table is it there. What group number is it? What period is it? What does a period mean? What is the atomic number? What does this tell you?

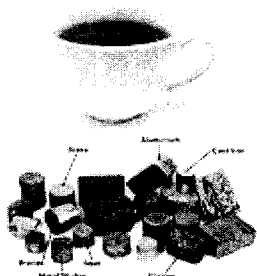


2. Find a compound. What is a compound? Name three compounds formulas that you know. Hint if you don't know one, use the periodic table to make one that makes sense.

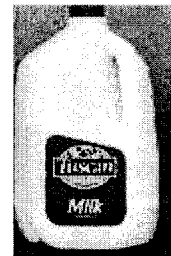


3. Make a heterogeneous mixture. What is a heterogeneous mixture? Name two ways you can immediately identify a heterogeneous mixture.

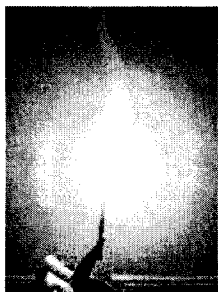
4. Make a homogeneous mixture. What is a homogeneous mixture? How can you tell if your mixture is homogeneous vs. heterogeneous?



5. Find a metal, non-metal, metalloid, and a noble gas. Explain all their properties.

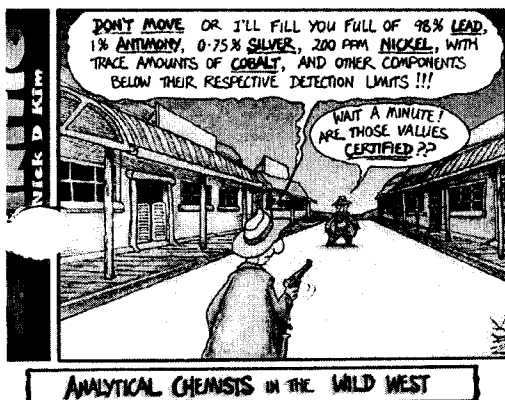


6. Construct a Magnesium ion ($^{24}\text{Mg}^{+2}$). Show Mr. Robinson. What is an ion? What is an isotope? Construct a Chlorine ion ($^{35}\text{Cl}^{-1}$). Discuss in your group...IS IT CORRECT?



Show Mr. Robinson

7. Create a physical change. Create a chemical change. Explain properties of a chemical change and a physical change. List 5 examples of each.



Chemistry Review Quiz

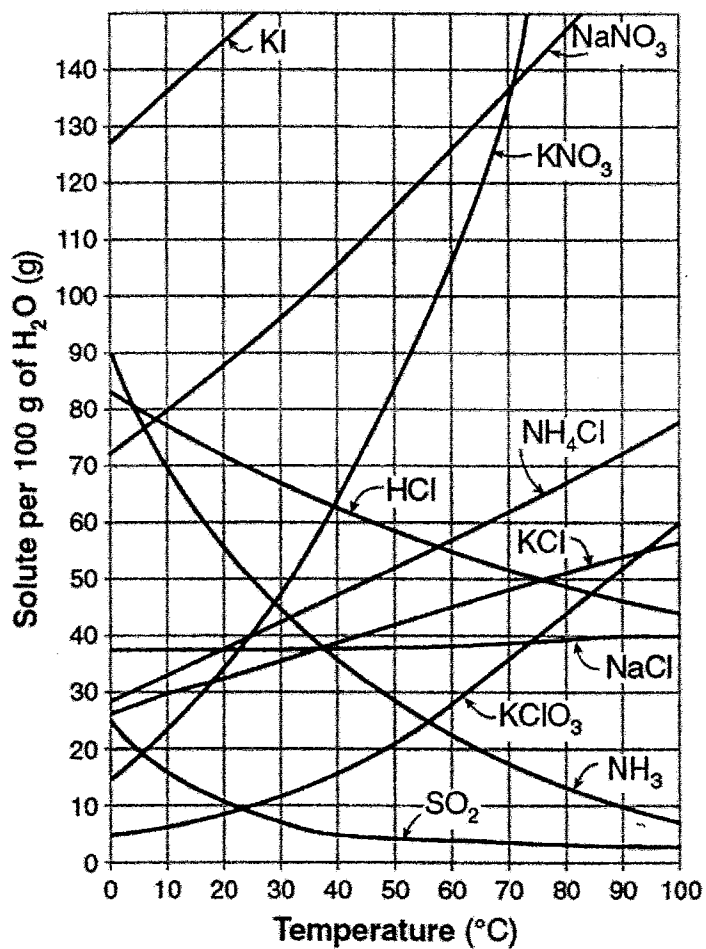
1. What is matter and what are the two different classifications?
2. An ice cube measuring 5.80 cm by 5.80 cm by 5.80 cm has a density of 0.917 g/mL. What is the mass?
3. The density of mercury is 13.6 g/mL. What is the volume of a 155-gram sample of mercury?
4. $\underline{\quad\quad} \text{Zn} + \underline{\quad\quad} \text{HCl} \rightarrow \underline{\quad\quad} \text{ZnCl}_2 + \underline{\quad\quad} \text{H}_2$
Type of Reaction _____
5. $\underline{\quad\quad} \text{H}_3\text{PO}_4 \rightarrow \underline{\quad\quad} \text{H}_4\text{P}_2\text{O}_7 + \underline{\quad\quad} \text{H}_2\text{O}$
Type of Reaction _____
6. $\underline{\quad\quad} \text{N}_2 + \underline{\quad\quad} \text{H}_2 \rightarrow \underline{\quad\quad} \text{NH}_3$
Type of Reaction _____
7. What do groups tell you on the periodic table?
8. What do periods tell you?

9. What does the Octet Rule say?

10. Explain the periodic trend atomic radius.

11. Draw the Lewis Dot structure or electron dot configuration for Selenium (Se).

Table G Solubility Curves



12. How much solute of NH₄Cl will dissolve at 90°C?

13. At what temperature will 110 grams of KNO₃ go into solution?